

2-[(*p*-Methoxyphenylcarbonyl)(1,2,4-triazol-1-yl)methyl]sulfanyl-4,6-dimethylpyrimidine

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Key indicators

Single-crystal X-ray study

$T = 293\text{ K}$

Mean $\sigma(\text{C}-\text{C}) = 0.003\text{ \AA}$

R factor = 0.045

wR factor = 0.125

Data-to-parameter ratio = 14.3

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

In the title compound, $\text{C}_{17}\text{H}_{17}\text{N}_5\text{O}_2\text{S}$, the dihedral angles between the plane of the 4,6-dimethyl-2-mercaptopyrimidine group and the plane of the triazole and *p*-methoxyphenylcarbonyl groups are $79.64(2)$ and $0.44(2)^\circ$, respectively. There are weak $\text{C}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{N}$ intermolecular interactions between the molecules in the crystal lattice.

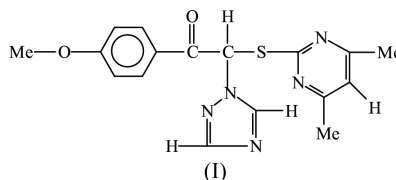
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Comment

As an important type of fungicides, triazole compounds are highly efficient, of low toxicity and capable of being absorbed through the stomach and intestines (Anderson, 1982; Shi *et al.*, 1995; Xu *et al.*, 2002). At present, the studies on triazole derivatives are mainly concentrated on compounds with triazole as the only active group. Reports of triazole compounds containing both a triazole group and another active group in a single molecule are rare. Some pyrimidines have been used as highly efficient fungicides of low toxicity (Tang & Li, 1998) in the control of powdery mildew. In this paper, we report the single-crystal structure of the title compound, (I).



In (I), the bond lengths and angles are generally normal in the phenyl ring and the triazole ring (Ji *et al.*, 2002; Liu *et al.*, 2002). The bond lengths and angles in the 4,6-dimethyl-2-mercaptopyrimidine group are in good agreement with an earlier report (Low *et al.*, 2002). Atom C9 of the central chain lies in the triazole ring (N1/N2/C11/N3/C10) plane, and the deviation from the least-squares plane through the ring atoms is smaller than $0.023(3)\text{ \AA}$. The atoms of the central chain (C5/C8/O2/C9/S1/C12) are almost planar, with a maximum displacement of $0.236(2)\text{ \AA}$ for C12. The dihedral angles formed by the pyrimidine, phenyl and triazole planes with the central chain (C5/C8/O2/C9/S1/C12) are $9.85(4)$, $10.29(6)$ and $80.70(8)^\circ$, respectively. The methoxy O1 and carbonyl C8 atoms lie in the C2–C7 phenyl plane, and the largest deviation from the least-squares plane through the ring atoms is $0.026(3)\text{ \AA}$. The dihedral angle between the triazole ring moiety and the phenyl ring is $79.22(2)^\circ$. Nine non-H atoms in the 4,6-dimethyl-2-mercaptopyrimidine group are also reasonably coplanar, and the largest deviation from the least-squares plane is $0.021(3)\text{ \AA}$. This plane is nearly parallel to the plane of the *p*-methoxyphenylcarbonyl group, with a dihedral

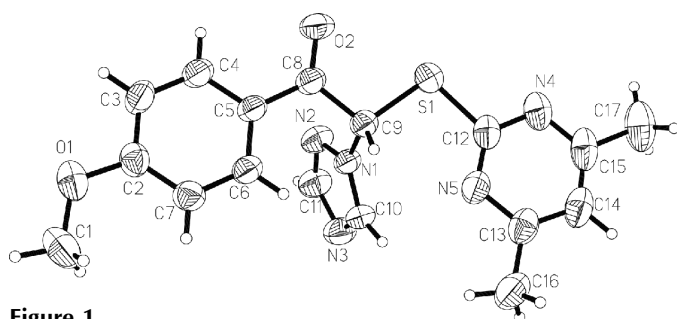


Figure 1
The structure of the title compound, showing 50% probability displacement ellipsoids and the atom-numbering scheme.

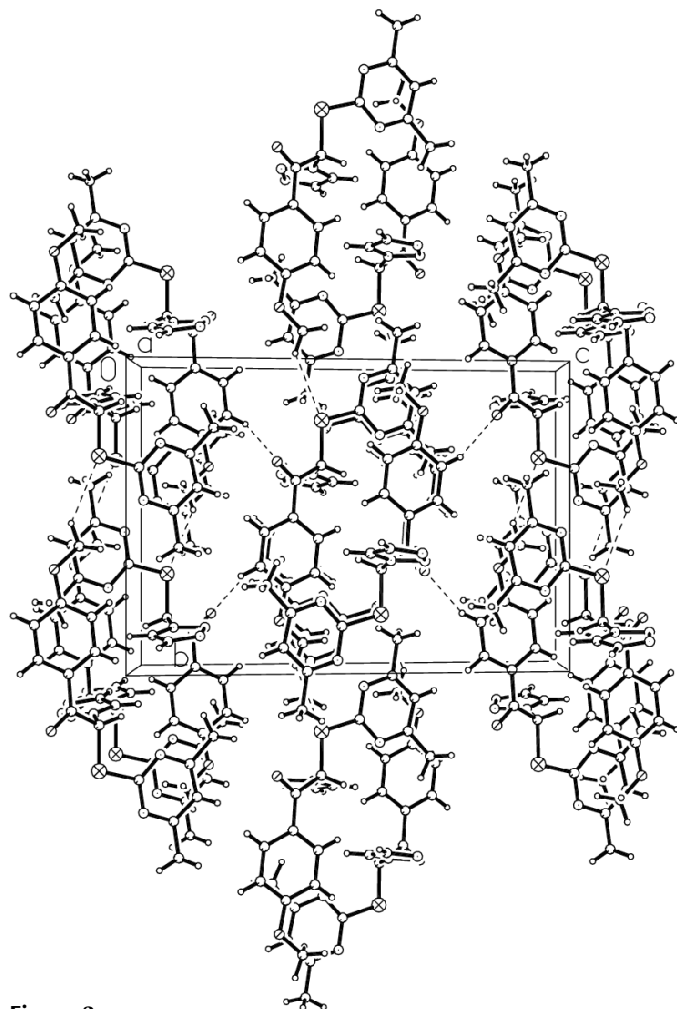


Figure 2
A view of the packing of the title compound.

angle of $0.44(2)^\circ$. The dihedral angle between the plane of the 4,6-dimethyl-2-mercaptopyrimidine group with the plane of the triazole moiety is $79.64(2)^\circ$.

Packing is stabilized by weak C—H \cdots O and C—H \cdots N interactions (Table 2) (Steiner, 1996; Jeffrey *et al.*, 1985).

Experimental

The title compound was prepared by reaction of (1,2,4-triazol-1-yl)(*p*-methoxyphenylcarbonyl)methane with 4,6-dimethyl-2-thioetherpyrimidine in chloroform. Single crystals of the title compound

suitable for X-ray measurements were obtained by recrystallization from ethyl ethanoate/cyclohexane ($v/v = 1:3$) at room temperature.

Crystal data

$C_{17}H_{17}N_5O_2S$
 $M_r = 355.42$
Monoclinic, $P2_1/c$
 $a = 8.0158(16) \text{ \AA}$
 $b = 12.462(3) \text{ \AA}$
 $c = 17.824(4) \text{ \AA}$
 $\beta = 99.89(3)^\circ$
 $V = 1754.0(7) \text{ \AA}^3$
 $Z = 4$

$D_x = 1.346 \text{ Mg m}^{-3}$
Mo $K\alpha$ radiation
Cell parameters from 20 reflections
 $\theta = 2-11^\circ$
 $\mu = 0.21 \text{ mm}^{-1}$
 $T = 293(2) \text{ K}$
Pillar, yellow
 $0.25 \times 0.20 \times 0.15 \text{ mm}$

Data collection

Oxford Instruments point-detector diffractometer
 $\theta/2\theta$ scans
Absorption correction: none
5553 measured reflections
3257 independent reflections
2334 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.031$

$\theta_{\text{max}} = 25.9^\circ$
 $h = -9 \rightarrow 9$
 $k = 0 \rightarrow 15$
 $l = -21 \rightarrow 21$
3 standard reflections
every 100 reflections
intensity decay: 0.9%

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.045$
 $wR(F^2) = 0.125$
 $S = 1.04$
3257 reflections
227 parameters
H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0696P)^2]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.25 \text{ e \AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.24 \text{ e \AA}^{-3}$
Extinction correction: *SHELXL97*
Extinction coefficient: 0.0062 (12)

Table 1

Selected geometric parameters (\AA).

S1—C12	1.775 (3)	O1—C1	1.429 (3)
S1—C9	1.816 (2)	O2—C8	1.215 (3)
O1—C2	1.364 (3)	N1—N2	1.368 (2)

Table 2

Hydrogen-bonding geometry (\AA , $^\circ$).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C3—H3B \cdots O2 ⁱ	0.93	2.52	3.409 (2)	159
C6—H6A \cdots N3 ⁱⁱ	0.93	2.57	3.340 (3)	139
C9—H9A \cdots N5	0.98	2.47	2.840 (3)	102
C10—H10A \cdots N5	0.93	2.58	3.054 (3)	111

Symmetry codes: (i) $2 - x, \frac{1}{2} + y, \frac{1}{2} - z$; (ii) $1 - x, 1 - y, 1 - z$.

The H atoms were fixed geometrically and were treated as riding on the parent C atoms, with C—H distances in the range 0.93–0.98 \AA . $U_{\text{iso}} = 1.2$ and 1.5 times U_{eq} of the parent atom.

Data collection: *R-Axis Software* (Rigaku, 1997); cell refinement: *R-Axis Software*; data reduction: *R-Axis Software*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXTL* (Sheldrick, 1990); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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